

Appendix C: Accuracy, Precision, and Uncertainty

How tall are you? How old are you? When you answered these everyday questions, you probably did it in round numbers such as "five foot, six inches" or "nineteen years, three months." But how true are these answers? Are you exactly 5' 6" tall? Probably not. You estimated your height at 5' 6" and just reported two significant figures. Typically, you round your height to the nearest inch, so that your actual height falls somewhere between 5' 5.5" and 5' 6.5" tall, or $5' 6" \pm 0.5"$. This $\pm 0.5"$ is the **uncertainty**, and it informs the reader of the precision of the **value** 5' 6".

What is uncertainty?

Whenever you measure something, there is always some uncertainty. There are two categories of uncertainty: **systematic** and **random**. Systematic uncertainties are those which consistently cause the value to be too large or too small. Systematic uncertainties include such things as reaction time, inaccurate meter sticks, optical parallax and miscalibrated balances. In principle, systematic uncertainties can be eliminated if you know they exist.

Random uncertainties are variations in the measurements which occur without a predictable pattern. If you make precise measurements, these uncertainties arise from the estimated part of the measurement. Random uncertainty can be reduced, but never eliminated. We need a technique to report the contribution of this uncertainty to the measured value.

How do I determine the uncertainty?

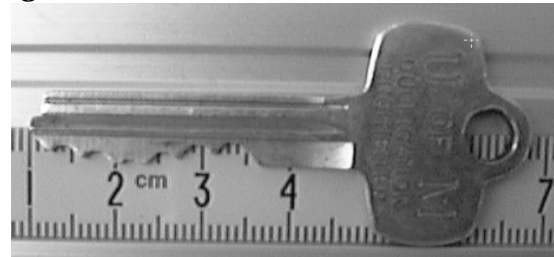
This Appendix will discuss two basic techniques for determining the uncertainty: **estimating the uncertainty** and measuring the **average deviation**. Which one you choose will depend on your need for precision. If you need a precise

determination of some value, the best technique is to measure that value several times and use the average deviation as the uncertainty. Examples of finding the average deviation are given below.

How do I estimate uncertainties?

If time or experimental constraints make repeated measurements impossible, then you will need to estimate the uncertainty. When you estimate uncertainties you are trying to account for anything that might cause the measured value to be different if you were to take the measurement again. For example, suppose you were trying to measure the length of a key, as in Figure C-1.

Figure C-1



If the true value was not as important as the magnitude of the value, you could say that the key's length was 5 cm give or take 1 cm. This is a crude estimate, but it may be acceptable. A better estimate of the key's length, as you saw in Appendix B, would be 5.81 cm. This tells us that the worst our measurement could be off is a fraction of a mm. To be more precise we estimate that fraction as about a third of a mm or that our value is 5.81 ± 0.03 cm.

How do I find the average deviation?

If estimating the uncertainty is not good enough for your situation, you can experimentally determine the uncertainty by making several measurements and calculating the average deviation of those measurements. To find the

average deviation: (1) Find the average of all your measurements; (2) Find the absolute value of the difference of each measurement from the average (deviations); (3) Find the average of all the deviations by adding them up and dividing by the number of measurements. Of course you need to take enough measurements to get a distribution for which the average has some meaning. Here are two examples of finding the average deviation.

In example 1, a class of six students was asked to find the mass of the same penny using the same balance. In example 2, another class measured a different penny using six different balances. Their results are listed below:

Class 1

Penny A massed by six different students on the same balance.

Mass (grams)
3.110
3.125
3.120
3.126
3.122
<u>3.120</u>
3.121 average
The magnitude of the deviations are: 0.011g, 0.004g, 0.001g, 0.005g, 0.001g, 0.001g
Sum of deviations: 0.023 g
Average deviation: (0.023g)/6 = 0.004 g
Mass of penny A: 3.121 ± 0.004 g

Class 2

Penny B massed by six different students on six different balances

Mass (grams)
3.140
3.133
3.144
3.118
3.132
<u>3.132</u>
3.133 average
The magnitude of the deviations are: 0.007g, 0.000g, 0.011g, 0.015g, 0.001g, 0.001g
Sum of deviations: 0.037 g

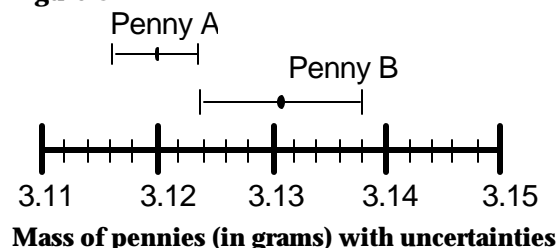
Average deviation: (0.039g)/6= 0.006 g
Mass of penny B: 3.131 ± 0.007 g

However you choose to determine the uncertainty, you should always clearly state your method in your report. For the remainder of this appendix, we will use the results of these two examples.

How do I know if two values are the same?

If we compare only the average masses of the two pennies we see that they are different. But now include the uncertainty in the masses. For penny A, the most likely mass is somewhere between 3.116g and 3.124g. For penny B, the most likely mass is somewhere between 3.124g and 3.138g. If you compare the ranges of the masses for the two pennies, as shown in Figure C-4, they just overlap. Given the uncertainty in the masses, we are able to conclude that the masses of the two pennies could be the same. If the range of the masses did not overlap, then we ought to conclude that the masses are probably different.

Figure C-4



Which result is more precise?

Using a meter stick, you might have an uncertainty of a fraction of a mm in measuring the length of a table or the width of a hair. Clearly, you are more certain about the length of the table than the width of the hair. Your measurement of the table is very precise and your measurement of the width of the hair is completely imprecise. To express this sense of precision, you need to calculate the percentage uncertainty. This is done by dividing the uncertainty by the value of the measurement and then multiplying by 100%. For example, class 1

and 2's precision can be calculated in the following way:

Precision of Class 1's value:

$$(0.004 \text{ g} \div 3.120 \text{ g}) \times 100\% = 0.1 \%$$

Precision of Class 2's value:

$$(0.007 \text{ g} \div 3.131 \text{ g}) \times 100\% = 0.2 \%$$

Class 1's results are more precise. This should not be surprising since class 2 introduced more uncertainty in their results by using six different balances instead of only one.

Which result is more accurate?

Accuracy is a measure of how your measured value compares with the real value. Imagine that class 2 made the measurement again using only one balance. Unfortunately, they chose a balance that was poorly calibrated. They analyzed their results and found the mass of penny B to be $3.556 \pm 0.004 \text{ g}$. This number is more precise than their previous result since the uncertainty is smaller, but the new measured value of mass is very different from their previous value.

We might conclude that this new value for the mass of penny B is different since the range of the new value does not overlap the range of the previous value. However, that conclusion would be **wrong** since our uncertainty has not taken into account the inaccuracy of the balance. To determine the accuracy of the measurement, we should check by measuring something that is known. This procedure is called calibration, and it is absolutely necessary for making accurate measurements.

Be cautious. It is possible to make measurements that are extremely precise and, at the same time, grossly inaccurate.

How can I do calculations with values that have uncertainty?

When you do calculations with values that have uncertainties, you will need to estimate (by calculation) the uncertainty in the result. There are mathematical techniques for doing this which depend on the statistical properties of your measurements. A very simple way to

estimate uncertainties this is to find the *largest possible uncertainty* the calculation could yield.

This will always overestimate the uncertainty of your calculation, but an overestimate is better than none at all. Examples of making calculations with uncertainties are shown below:

Addition:

$$3.120 \pm 0.004 \text{ g} + 3.131 \pm 0.007 \text{ g} = ?$$

First find the sum of the values:

$$3.120 \text{ g} + 3.131 \text{ g} = 6.251 \text{ g}$$

Next find the largest possible value:

$$3.124 \text{ g} + 3.138 \text{ g} = 6.262 \text{ g}$$

The uncertainty is the difference between the two:

$$6.262 \text{ g} - 6.251 \text{ g} = 0.011 \text{ g}$$

This is the same as just adding the uncertainties:

$$0.004 \text{ g} + 0.007 \text{ g} = 0.011 \text{ g}$$

Subtraction:

$$3.131 \pm 0.007 \text{ g} - 3.120 \pm 0.004 \text{ g} = ?$$

First find the differences of the values:

$$3.131 \text{ g} - 3.120 \text{ g} = 0.011 \text{ g}$$

Next find the largest possible uncertainty:

$$0.007 \text{ g} + 0.004 \text{ g} = 0.011 \text{ g}$$

So the result is $0.011 \pm 0.011 \text{ g}$.

Notice that zero is included in the range, so it is possible that there is no difference in the masses of the pennies, as we saw before.

Multiplication:

$$3.131 \pm 0.013 \text{ g} \times 6.1 \pm 0.2 \text{ cm} = ?$$

First multiply the values:

$$3.131 \text{ g} \times 6.1 \text{ cm} = 19.1 \text{ g-cm}$$

Next find the maximum possible value (or the smallest):

$$3.144 \text{ g} \times 6.3 \text{ cm} = 19.8 \text{ g-cm}$$

And the difference gives the uncertainty:

$$19.8 \text{ g-cm} - 19.1 \text{ g-cm} = 0.7 \text{ g-cm}$$

So the result is $19.1 \pm 0.7 \text{ g-cm}$.

Division:

$$3.131 \pm 0.013 \text{ g} \div 3.120 \pm 0.008 \text{ g} = ?$$

First divide the values:

$$3.131 \div 3.120 = 1.0035$$

Next, find the maximum possible value (or the smallest):

$$3.144\text{g} \div 3.112\text{g} = 1.0103.$$

(Notice the largest possible value for the numerator and the smallest possible value for the denominator gives the largest result.)

And the deviation of this value:

$$1.0103 - 1.0035 = 0.0068$$

So the result is 1.004 ± 0.007

The same ideas can be carried out with more complicated calculations. Remember this will always give you an overestimate of your

uncertainty. Other calculational techniques give better estimates for uncertainties. If you wish to use them, please discuss it with your instructor to see if they are appropriate.

These techniques help you estimate the random uncertainty that always occurs in measurements. They will not help account for mistakes or poor measurement procedures. There is no substitute for taking data with the utmost of care. A little forethought about the possible sources of uncertainty can go a long way in ensuring precise and accurate data.

PRACTICE EXERCISES:

C-1. Consider the following results for different experiments. Determine if they agree with the accepted result listed to the right. Also calculate the precision for each result.

a) $g = 10.4 \pm 1.1 \text{ m/s}^2$

$g = 9.8 \text{ m/s}^2$

b) $T = 1.5 \pm 0.1 \text{ sec}$

$T = 1.1 \text{ sec}$

c) $k = 1368 \pm 45 \text{ N/m}$

$k = 1300 \pm 50 \text{ N/m}$

Answer: a) Yes, 11%; b) No, 7%; c) Yes, 3.3%

C-2. The area of a rectangular metal plate was found by measuring its length and its width. The length was found to be $5.37 \pm 0.05 \text{ cm}$. The width was found to be $3.42 \pm 0.02 \text{ cm}$. What is the area and the average deviation?

Answer: $18.4 \pm 0.3 \text{ cm}^2$

C-3. Each member of your lab group weighs the cart and two mass sets twice. The following table shows this data. Calculate the total mass of the cart with each set of masses and for the two sets of masses combined.

Cart (grams)	Mass set 1 (grams)	Mass set 2 (grams)
201.3	98.7	95.6
201.5	98.8	95.3
202.3	96.9	96.4
202.1	97.1	96.2
199.8	98.4	95.8
200.0	98.6	95.6

Answer: Cart and set 1: $299.3 \pm 2.2 \text{ g}$.
 Cart and set 2: $297.0 \pm 1.8 \text{ g}$.
 Cart and both sets: $395.1 \pm 2.$

